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Chemical Kinetics Ncert Solutions

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Solution: The reaction, X → Y, follows second order kinetics hence the rate law equation will be. Rate = kC², where C = [X] If concentration of X increases three times, now, [X] = 3C mol L⁻¹. ∴ Rate = k [3C]² = 9kC². Thus the rate of reaction will become 9 times.

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the branch of chemistry, which deals with the study of reaction rates and their mechanisms, called chemical kinetics. The word kinetics is derived from the Greek word 'kinesis' meaning movement. Thermodynamics tells only about the feasibility of a reaction whereas chemical kinetics tells about the rate of a reaction. For example ...

Chemical Kinetics - NCERT

Chapter 4: Chemical Kinetics. Chemical Kinetics – The branch of chemistry which deals with the study of reaction rates and their mechanism. The rate of change of concentration of any of the reactant or product with time at any particular moment of time is known as Rate of reaction. NCERT Exercises Solutions Chapter 4

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