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Holt Modern Chemistry: Problem-Solving Workbook 1st Edition

Holt ChemFile: Problem-Solving Workbook 21 Significant Figures Name Class Date Problem Solving continued Sample Problem 2 In an experiment to identify an unknown gas, it is found that 1.82 L of the gas has a mass of 5.430 g. What is the density of the gas in g/L? Solution ANALYZE What is given in the problem? the measured mass and volume of the gas

Skills Worksheet Problem Solving

Holt ChemFile: Problem-Solving Workbook 58 Mole Concept Name Class Date Problem Solving continued CONVERTING NUMBER OF ATOMS OF AN ELEMENT TO MASS Sample Problem 4

uses the progression of steps 1→2→3 to convert from the mass of an element to the number of atoms. In order to calculate the mass from a given number of atoms, these steps will be reversed.

Skills Worksheet Problem Solving

Holt ChemFile: Problem-Solving Workbook 272 Titrations Titrations Chemists have many methods for determining the quantity of a substance present in a solution or other mixture. One common method is titration, in which a solu-tion of known concentration reacts with a sample containing the substance of unknown quantity.

Skills Worksheet Problem Solving

Holt ChemFile: Problem-Solving Workbook 1 Conversions Conversions One of the aims of chemistry is to describe changes—to tell what changed, how it changed, and what it changed into. Another aim of chemistry is to look at matter and its changes and to ask questions such as how much, how big,

how hot, how many, how hard, and how long did it take.

Skills Worksheet Problem Solving

Holt ChemFile: Problem-Solving Workbook 49 Mole Concept Name Class Date • Problem Solving continued PROBLEMS INVOLVING ATOMS AND ELEMENTS Sample Problem 1 A chemist has a jar containing 388.2 g of iron filings. How many moles of iron does the jar contain? Solution ANALYZE What is given in the problem? mass of iron in grams

continued - PC\|MAC

Holt ChemFile: Problem-Solving Workbook 261 pH Name Class Date Problem Solving continued Sample Problem 1 A HCl solution has a concentration of 0.0050 M. Calculate [OH] and [H 3O] for this solution. HCl is a strong acid, so assume it is 100% ionized. Solution ANALYZE What is given in the problem? the molarity of the HCl solution, and the fact

Skills Worksheet Problem Solving

Holt ChemFile: Problem-Solving Workbook 99 Stoichiometry Name Class Date Problem Solving continued Sample Problem 1 Ammonia is made industrially by reacting nitrogen and hydrogen under pressure, at high temperature, and in the presence of a catalyst. The equation is N 2(g) 3H 2(g)→2NH 3(g). If 4.0 mol of H 2 react, how many moles of NH 3 will be produced?

Skills Worksheet Problem Solving

Holt ChemFile: Problem-Solving Workbook97Stoichiometry. Stoichiometry. So far in your chemistry course, you have learned that chemists count quantities of elements and compounds in terms of moles and that they relate moles of a substance to mass by using the molar mass.

Skills Worksheet Problem Solving - Mole Cafe

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Workbook 1 Conversions Conversions One of the aims of chemistry is to describe changes—to tell what changed, how it changed, and what it changed into. Another aim of chemistry is to look at matter and its changes and to ask questions such as how much, how big, how hot, how many, how hard, and how long did it take.

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Holt ChemFile: Problem-Solving Workbook 201 Concentration of Solutions Sample Problem 3 Determine the molal concentration of a solution containing 81.3 g of ethylene glycol, HOCH 2 CH 2 OH, dissolved in 166 g of water. Solution ANALYZE What is given in the problem? the mass of ethylene glycol dissolved, and What are you asked to find? Items Data

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Holt ChemFile: Problem-Solving Workbook 57 Mole Concept Name Class Date Problem Solving continued Is the answer reasonable? Yes; 2 g of boron is about 1/5 of the molar mass of boron. Therefore, 2.00 g boron will contain about 1/5 of an Avogadro's constant of atoms. Practice 1.

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Holt ChemFile: Problem-Solving Workbook 98 Stoichiometry Name Class Date Problem Solving continued General Plan for Solving Stoichiometry Problems Convert using the mole ratio A, given in the balanced chemical equation Mass of substance A Amount in mol of substance A Amount in mol of substance B Convert using the molar mass of A 1

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